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Surname

Other names

**Pearson Edexcel**  
**International**  
**Advanced Level**

Centre Number

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Candidate Number

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# Chemistry

**Advanced Subsidiary**

**Unit 2: Application of Core Principles of Chemistry**

Friday 9 June 2017 – Afternoon

**Time: 1 hour 30 minutes**

Paper Reference

**WCH02/01**

**Candidates may use a calculator.**

Total Marks

## Instructions

- Use **black** ink or **black** ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided  
– *there may be more space than you need.*

## Information

- The total mark for this paper is 80
- The marks for **each** question are shown in brackets  
– *use this as a guide as to how much time to spend on each question.*
- Questions labelled with an **asterisk** (\*) are ones where the quality of your written communication will be assessed  
– *you should take particular care with your spelling, punctuation and grammar, as well as the clarity of expression, on these questions.*
- A Periodic Table is printed on the back cover of this paper.

## Advice

- Read each question carefully before you start to answer it.
- Keep an eye on the time.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ►

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Pearson

## SECTION A

Answer ALL the questions in this section. You should aim to spend no more than 20 minutes on this section. For each question, select one answer from A to D and put a cross . If you change your mind, put a line through the box  and then mark your new answer with a cross .

1 Which is the shortest covalent bond?

- A H—H
- B H—N
- C H—S
- D H—Br

(Total for Question 1 = 1 mark)

2 Which compound contains a bond with the **greatest** polarity?

- A Ammonia, NH<sub>3</sub>
- B Hydrogen fluoride, HF
- C Methane, CH<sub>4</sub>
- D Water, H<sub>2</sub>O

(Total for Question 2 = 1 mark)

3 Which compound has polar bonds but non-polar molecules?

- A Carbon monoxide, CO
- B Hydrogen sulfide, H<sub>2</sub>S
- C Phosphorus(III) chloride, PCl<sub>3</sub>
- D Tetrafluoromethane, CF<sub>4</sub>

(Total for Question 3 = 1 mark)

4 Cyclohexane is a non-polar liquid. Therefore

- A sodium chloride is very soluble in cyclohexane.
- B cyclohexane conducts electricity.
- C a jet of cyclohexane is deflected by a charged rod.
- D cyclohexane forms two layers when mixed with water.

(Total for Question 4 = 1 mark)

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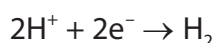


5 In which reaction is calcium oxidised?

- A  $\text{Ca} + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{H}_2$
- B  $\text{CaO} + 2\text{K} \rightarrow \text{Ca} + \text{K}_2\text{O}$
- C  $\text{CaO} + \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2$
- D  $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$

(Total for Question 5 = 1 mark)

6 Consider the following ionic half-equations



When these ionic half-equations are combined, the full ionic equation is

- A  $\text{Al} + 2\text{H}^+ \rightarrow \text{Al}^{3+} + \text{H}_2$
- B  $\text{Al} + 2\text{H}^+ + 2\text{e}^- \rightarrow \text{Al}^{3+} + \text{H}_2 + 3\text{e}^-$
- C  $\text{Al} + 6\text{H}^+ \rightarrow \text{Al}^{3+} + 3\text{H}_2$
- D  $2\text{Al} + 6\text{H}^+ \rightarrow 2\text{Al}^{3+} + 3\text{H}_2$

(Total for Question 6 = 1 mark)

7 The metal salt which gives a red colour in a flame test is

- A barium nitrate.
- B lithium chloride.
- C potassium nitrate.
- D sodium chloride.

(Total for Question 7 = 1 mark)

Use this space for any rough working. Anything you write in this space will gain no credit.



- 8 What is the trend in the thermal stability of the carbonates and nitrates as Group 2 is descended?

	Carbonates	Nitrates
<input type="checkbox"/> A	decreases	decreases
<input type="checkbox"/> B	decreases	increases
<input type="checkbox"/> C	increases	decreases
<input type="checkbox"/> D	increases	increases

(Total for Question 8 = 1 mark)

- 9 Which pair of compounds has the more soluble hydroxide and the more soluble sulfate?

- A  $\text{Mg}(\text{OH})_2$  and  $\text{MgSO}_4$
- B  $\text{Mg}(\text{OH})_2$  and  $\text{SrSO}_4$
- C  $\text{Sr}(\text{OH})_2$  and  $\text{MgSO}_4$
- D  $\text{Sr}(\text{OH})_2$  and  $\text{SrSO}_4$

(Total for Question 9 = 1 mark)

- 10 The table shows the measurement uncertainty of each reading for some laboratory apparatus.

Laboratory apparatus	Measurement uncertainty of each reading / $\text{cm}^3$
burette	$\pm 0.05$
measuring cylinder, $25 \text{ cm}^3$	$\pm 0.5$
pipette, $25 \text{ cm}^3$	$\pm 0.06$
volumetric flask, $25 \text{ cm}^3$	$\pm 0.1$

The item of laboratory apparatus that would measure a volume of  $25 \text{ cm}^3$  with the **lowest** percentage uncertainty is the

- A burette.
- B measuring cylinder,  $25 \text{ cm}^3$ .
- C pipette,  $25 \text{ cm}^3$ .
- D volumetric flask,  $25 \text{ cm}^3$ .

(Total for Question 10 = 1 mark)



11 On adding chlorine water to aqueous sodium bromide, the resulting solution is

- A colourless.
- B pale yellow-green.
- C red-brown.
- D purple.

(Total for Question 11 = 1 mark)

12 A solid silver halide was tested as follows:

Test	Result
action of sunlight	solid turned grey
addition of dilute ammonia	solid did not dissolve
addition of concentrated ammonia	solid dissolved

The silver halide is

- A AgF
- B AgCl
- C AgBr
- D AgI

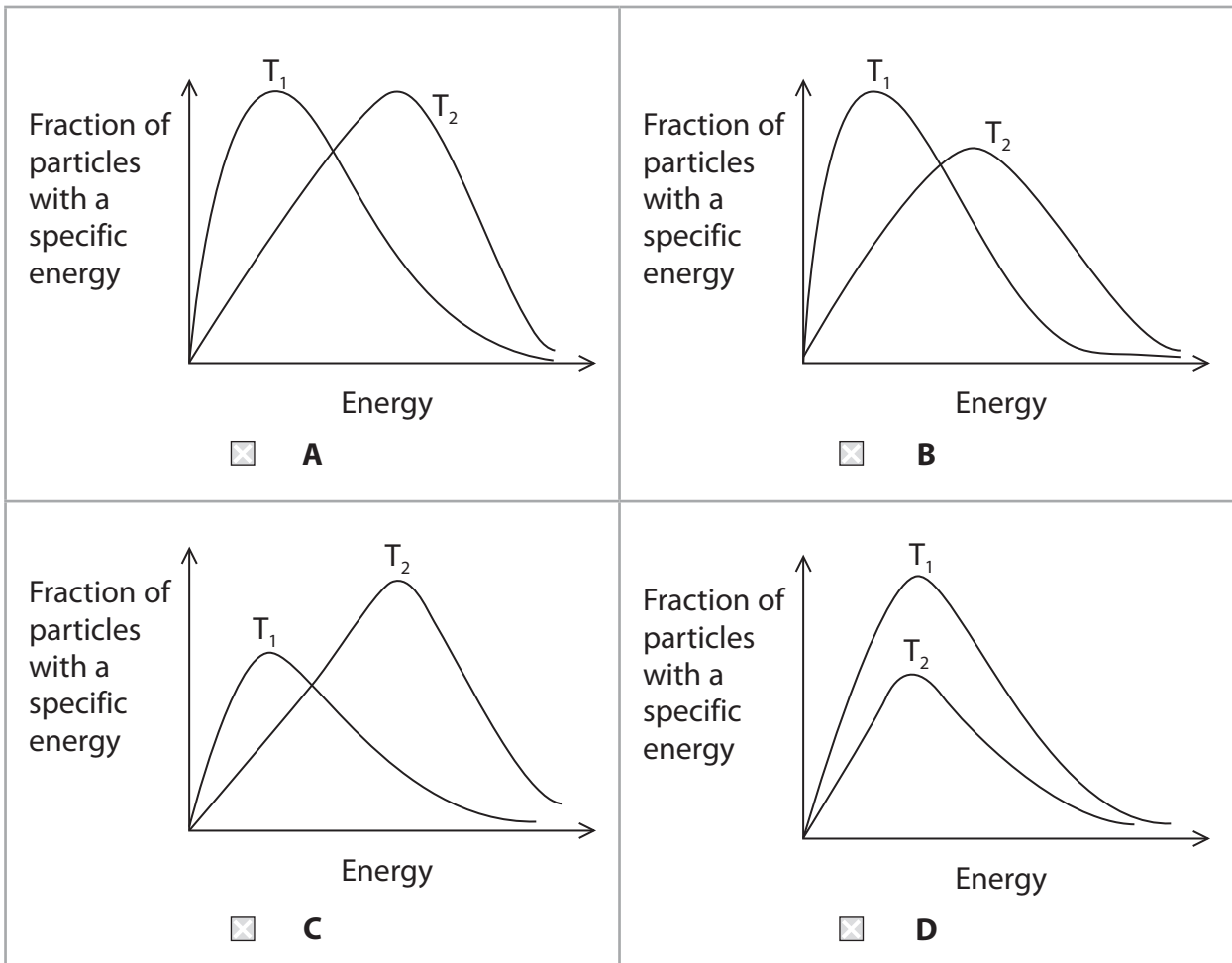
(Total for Question 12 = 1 mark)

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13 This question is about the Maxwell-Boltzmann energy distribution.

The diagram for an **increase** in temperature from  $T_1$  to  $T_2$  is



(Total for Question 13 = 1 mark)

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14 The rate of the reaction between calcium carbonate and acid **increases** when

- A the particle size of the calcium carbonate decreases.
- B  $1 \text{ mol dm}^{-3}$  nitric acid is used instead of  $1 \text{ mol dm}^{-3}$  hydrochloric acid.
- C  $0.5 \text{ mol dm}^{-3}$  sulfuric acid is used instead of  $1 \text{ mol dm}^{-3}$  hydrochloric acid.
- D the pressure is increased.

(Total for Question 14 = 1 mark)

15 The most significant factor determining the trend in the rate of hydrolysis of halogenobutanes is

- A the electronegativity of the halogen.
- B the magnitude of the halogen ionisation energy.
- C the oxidising ability of the halogen.
- D the carbon-halogen bond strength.

(Total for Question 15 = 1 mark)

16 The action of ultraviolet radiation on an oxygen molecule high in the atmosphere results in

- A no change because  $\text{O}_2$  has no dipole.
- B only increased bond vibration.
- C the production of two oxygen atoms.
- D the formation of an oxide ion.

(Total for Question 16 = 1 mark)

17 The mass spectrum of propanal can be clearly distinguished from the mass spectrum of propanone. Only the propanal spectrum has a large peak due to the

- A  $\text{C}_3\text{H}_6\text{O}^+$ , molecular ion,  $m/e = 58$
- B  $\text{C}_3\text{H}_5\text{O}^+$  fragment,  $m/e = 57$
- C  $\text{C}_2\text{H}_5^+$  fragment,  $m/e = 29$
- D  $\text{CH}_3^+$  fragment,  $m/e = 15$

(Total for Question 17 = 1 mark)

Use this space for any rough working. Anything you write in this space will gain no credit.



18 A sample of butan-2-ol was oxidised by heating under reflux with an oxidising agent and then the product was separated for infrared analysis. Apart from the peaks due to the C—C and C—H bonds, which peaks would be present in the IR spectrum of the oxidation product?

- A A peak due to C=O only.
- B A peak due to O—H only.
- C Peaks due to C=O and O—H.
- D Peaks due to C—O, C=O and O—H.

(Total for Question 18 = 1 mark)

19 Which greenhouse gas is produced **only** as a result of anthropogenic activity?

- A carbon dioxide
- B dichlorodifluoromethane
- C methane
- D water vapour

(Total for Question 19 = 1 mark)

20 The first ionisation energy of strontium is less endothermic than that of calcium.

The best explanation for this is that strontium has

- A more protons.
- B more protons and neutrons.
- C 18 and not 8 electrons in its outer shell.
- D more inner electron shells.

(Total for Question 20 = 1 mark)

**TOTAL FOR SECTION A = 20 MARKS**





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## SECTION B

Answer ALL the questions. Write your answers in the spaces provided.

- 21 Tablets of potassium iodate(V),  $\text{KIO}_3$ , may be used to protect against the build-up of radioactive iodine in the body. The use of potassium iodate(V) is preferred to potassium iodide because, in hot and humid conditions, the potassium iodate(V) can be stored for much longer.

A very old sample of potassium iodate(V) tablets, which originally contained 85 mg of  $\text{KIO}_3$  per tablet, was analysed using the following procedure.

A tablet was crushed, dissolved in deionised water and the solution and washings added to a conical flask. Then potassium iodide, KI, and hydrochloric acid, both in excess, were added to the conical flask. This mixture was titrated with  $0.0600 \text{ mol dm}^{-3}$  sodium thiosulfate solution.

This procedure was repeated and the following burette readings were obtained.

Titration	1	2	3
Final volume / $\text{cm}^3$	19.90	39.70	39.85
Initial volume / $\text{cm}^3$	0.00	19.90	20.00
Volume added / $\text{cm}^3$	19.90	19.80	19.85
Mean titre / $\text{cm}^3$	19.85		

- (a) State why it was **not** essential to carry out the third titration.

(1)

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- (b) Starch was added to the titration mixture in order to make the end-point easier to observe.

- (i) State the colour change observed at the end-point with starch.

(1)

From ..... to .....

- (ii) Identify the substance in the titration mixture that reacts with starch.

(1)

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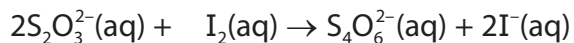


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(c) The equations for the reactions involved are



(i) Calculate the number of moles of sodium thiosulfate that reacted. (1)

(ii) Calculate the number of moles of iodine that reacted with the thiosulfate. (1)

(iii) Calculate the mass in **milligrams** of potassium iodate(V) in each tablet. Give your answer to **three** significant figures. (3)



(iv) In a radiation emergency, the recommended adult dose is 170 mg of  $\text{KIO}_3$  every 24 hours.

Using your result to (c)(iii), suggest whether or not the old tablets of potassium iodate(V) are suitable for use. Justify your answer.

(2)

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(v) The experiment was repeated with a different batch of tablets. The conical flask contained  $2.15 \times 10^{-4}$  mol of potassium iodate(V).

Calculate the minimum volume of  $0.100 \text{ mol dm}^{-3}$  hydrochloric acid that should be added to ensure that all of the potassium iodate(V) is converted to iodine and hence suggest an appropriate volume to use.

(3)

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(d) Potassium iodate(V) can be produced from iodine and potassium hydroxide.

- (i) Give the oxidation numbers of iodine in the iodine-containing species in the following equation. Hence classify the reaction.

(2)



Oxidation  
Number

.....

Type of reaction.....

- (ii) State the conditions necessary for this reaction to occur.

(1)

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**(Total for Question 21 = 16 marks)**



22 This is a question about alcohols.

(a) There are two alcohol structural isomers with the molecular formula,  $C_3H_8O$ .

Give the **skeletal** formula of these isomers, their systematic names and the classification of the type of alcohol in each case.

(3)

Skeletal formula	Name	Classification

(b) Ethanol can be oxidised by acidified sodium dichromate(VI) to ethanal and then to ethanoic acid. The apparatus may be set up in two ways.

Reflux apparatus	Distillation apparatus
<p>Labels: open top, water out, Liebig condenser, water in, pear-shaped flask, anti-bumping granules, HEAT</p>	<p>Labels: thermometer, Liebig condenser, distillate, anti-bumping granules, round bottomed flask, HEAT</p>



- (i) Complete the ionic half-equation for the reduction of the dichromate(VI) ions to chromium(III) ions. State symbols are not required.

(2)



- (ii) Describe how the reflux apparatus ensures that any ethanal initially produced is further oxidised to ethanoic acid.

(1)

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- (iii) The distillation apparatus effectively separates ethanal from ethanol because of the large difference in boiling temperatures, which is a result of the hydrogen bonding between the molecules in ethanol.

Compound	Boiling temperature / °C
Ethanol, CH <sub>3</sub> CH <sub>2</sub> OH	79
Ethanal, CH <sub>3</sub> CHO	21

Draw a hydrogen bond between two ethanol molecules. Clearly indicate any relevant dipoles and lone pairs of electrons. Label the bond angle about the hydrogen involved in the hydrogen bond and give its value.

(3)

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(iv) Explain why hydrogen bonds do **not** form between ethanal molecules.

(1)

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(c) Alcohols can be converted into halogenoalkanes.

(i) Write the equation for the reaction between methanol,  $\text{CH}_3\text{OH}$ , and phosphorus(V) chloride,  $\text{PCl}_5$ .

(1)

(ii) State the experimental observation from this reaction.

(1)

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\*(iii) Chloroethane can be made from a mixture of ethanol, potassium chloride and concentrated sulfuric acid. Explain why chloroethane can be made in this way, but iodoethane cannot be made from a similar mixture using potassium iodide instead of potassium chloride.

You may use equations to support your explanation.

(3)

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(d) Alcohols can be produced from the reaction of halogenoalkanes with aqueous alkali.

- (i) Draw the mechanism for this reaction with 1-bromopropane. Show the lone pair involved in the mechanism and any relevant dipoles and curly arrows.

(3)

- (ii) The reaction of 1-bromopropane with concentrated alcoholic alkali forms a different organic product. Name the type of reaction and give the **displayed** formula of the product.

(2)

Name of reaction.....

Displayed formula of product



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(e) How would you test for the OH group in 2-methylpropan-2-ol without using phosphorus(V) chloride?

Name the reagent and state the observation for a positive test.

(2)

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**(Total for Question 22 = 22 marks)**

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**TOTAL FOR SECTION B = 38 MARKS**



## SECTION C

Answer ALL the questions. Write your answers in the spaces provided.

- 23 Boron nitride, BN, is a compound first made commercially in the 1940s from boric acid and ammonia, in an atmosphere of nitrogen.

It forms structures analogous to graphite and diamond because it is isoelectronic with these corresponding carbon structures. Boron nitride has also been used to form nanotube structures in a similar way to carbon.

Just as synthetic diamonds are produced from graphite by using high temperatures and high pressures, the diamond-like cubic boron nitride can also be made from heating the graphite-like hexagonal boron nitride under high pressure.

Boron nitride forms ceramic materials with very high thermal and chemical stability and, a wide range of uses. For example, they are stable in air up to  $1000^{\circ}\text{C}$ , which is an advantage over similar graphite materials. The hexagonal form of boron nitride is a very effective lubricant and is also used in cosmetics. However, it is an electrical insulator, in contrast to graphite, which is a good electrical conductor.

- (a) (i) Write the equation for the formation of boron nitride from boric acid,  $\text{H}_3\text{BO}_3$ , and ammonia.

State symbols are not required.

(1)

- (ii) Suggest why this reaction is carried out in an atmosphere of nitrogen.

(1)

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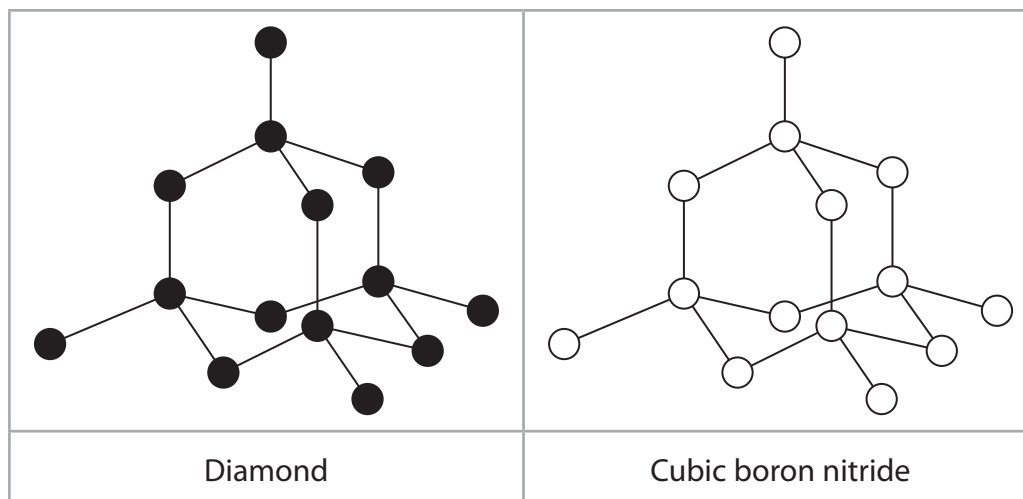


(b) The structure of the cubic boron nitride corresponds to the diamond structure. The boron and nitrogen atoms alternate throughout the structure.

- (i) In the left hand box, the diagram shows a section of the diamond structure, where each black circle represents a carbon atom.

In the right hand box label all the nitrogen and boron atoms in the diagram of cubic boron nitride.

(1)



- (ii) State the bond angle and shape around the carbon atoms in diamond and fully justify your answer.

(4)

Bond angle ..... Shape .....

Justification .....

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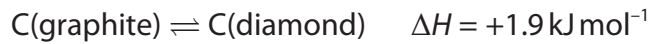


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(c) The equilibrium between graphite and diamond is



The density of graphite is  $2.27 \text{ g cm}^{-3}$  and the density of diamond is  $3.51 \text{ g cm}^{-3}$ .

\* (i) Suggest why a very high temperature and high pressure are needed to convert graphite to diamond.

(4)

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(ii) The use of a catalyst in the conversion of graphite to diamond has been reported. Describe how the addition of a catalyst can lower the temperature required for a reaction.

(3)

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- (d) Diamond and graphite are stable in air up to approximately 800°C. Identify **one** of the products if diamond or graphite is heated in air above this temperature. (1)

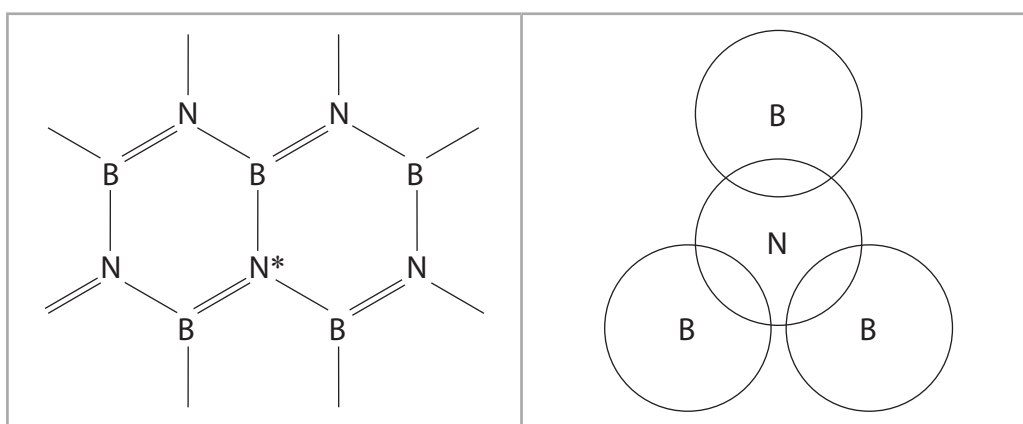
(e) The structure of hexagonal boron nitride corresponds to that of graphite.

- (i) The simplified diagram in the left hand box shows the bonding in hexagonal boron nitride.

In the right hand box, complete the dot and cross diagram showing only the electrons around the nitrogen atom which is labelled with an asterisk (\*).

Use (×) for the nitrogen electrons and (•) for the boron electrons.

(1)



- \*(ii) Describe how each carbon atom is bonded in the graphite structure and hence explain why graphite is a good conductor of electricity. Suggest why hexagonal boron nitride is an electrical insulator.

(3)



(iii) Graphite and the hexagonal boron nitride are both used as lubricants because of the weak intermolecular forces between the layers of hexagonal rings. Identify these intermolecular forces and describe how they arise.

(3)

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**(Total for Question 23 = 22 marks)**

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**TOTAL FOR SECTION C = 22 MARKS**  
**TOTAL FOR PAPER = 80 MARKS**

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# The Periodic Table of Elements

	1	2	Key										3	4	5	6	7	0 (8)
(1)	(2)	(3)	(4)	(5)	(6)	(7)	(8)	(9)	(10)	(11)	(12)	(13)	(14)	(15)	(16)	(17)	(18)	
		relative atomic mass																
		atomic symbol																
		name																
		atomic (proton) number																
6.9	9.0	45.0	47.9	50.9	52.0	54.9	55.8	58.9	58.7	63.5	65.4	10.8	12.0	14.0	16.0	19.0	4.0	
Li	Be	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	B	C	N	O	F	He	
lithium	beryllium	scandium	titanium	vanadium	chromium	manganese	iron	cobalt	nickel	copper	zinc	boron	carbon	nitrogen	oxygen	fluorine	helium	
3	4	21	22	23	24	25	26	27	28	29	30	5	6	7	8	9	2	
23.0	24.3	40.1	47.9	50.9	52.0	54.9	55.8	58.9	58.7	63.5	65.4	27.0	28.1	31.0	32.1	35.5	39.9	
Na	Mg	Ca	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Al	Si	P	S	Cl	Ar	
sodium	magnesium	calcium	titanium	vanadium	chromium	manganese	iron	cobalt	nickel	copper	zinc	aluminium	silicon	phosphorus	sulfur	chlorine	argon	
11	12	20	22	23	24	25	26	27	28	29	30	13	14	15	16	17	18	
39.1	87.6	88.9	91.2	92.9	95.9	[98]	101.1	102.9	106.4	107.9	112.4	69.7	72.6	74.9	79.0	79.9	83.8	
K	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	Ga	Ge	As	Se	Br	Kr	
potassium	strontium	yttrium	zirconium	niobium	molybdenum	technetium	ruthenium	rhodium	palladium	silver	cadmium	gallium	germanium	arsenic	selenium	bromine	krypton	
19	38	39	40	41	42	43	44	45	46	47	48	31	32	33	34	35	36	
85.5	87.6	88.9	91.2	92.9	95.9	[98]	101.1	102.9	106.4	107.9	112.4	114.8	118.7	121.8	127.6	126.9	131.3	
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe	
rubidium	strontium	yttrium	zirconium	niobium	molybdenum	technetium	ruthenium	rhodium	palladium	silver	cadmium	indium	tin	antimony	tellurium	iodine	xenon	
37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54	
132.9	137.3	138.9	178.5	180.9	183.8	186.2	190.2	192.2	195.1	197.0	200.6	204.4	207.2	209.0	[209]	[210]	[222]	
Cs	Ba	La*	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn	
caesium	barium	lanthanum	hafnium	tantalum	tungsten	rhenium	osmium	iridium	platinum	gold	mercury	thallium	lead	bismuth	polonium	astatine	radon	
55	56	57	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86	
[223]	[226]	[227]	[261]	[262]	[266]	[264]	[277]	[268]	[271]	[272]								
Fr	Ra	Ac*	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg								
francium	radium	actinium	rutherfordium	dubnium	seaborgium	bohrium	hassium	meitnerium	darmstadtium	roentgenium								
87	88	89	104	105	106	107	108	109	110	111								

Elements with atomic numbers 112-116 have been reported but not fully authenticated

140	141	144	150	152	157	159	163	165	167	169	173	175
Ce	Pr	Nd	Pm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu
cerium	praseodymium	neodymium	promethium	europium	gadolinium	terbium	dysprosium	holmium	erbium	thulium	ytterbium	lutetium
58	59	60	61	62	64	65	66	67	68	69	70	71
232	[231]	238	[237]	[242]	[247]	[245]	[251]	[254]	[253]	[256]	[254]	[257]
Th	Pa	U	Np	Pu	Cm	Bk	Cf	Es	Fm	Md	No	Lr
thorium	protactinium	uranium	neptunium	plutonium	curium	berkelium	californium	einsteinium	fermium	mendeleevium	nobelium	lawrencium
90	91	92	93	94	96	97	98	99	100	101	102	103

\* Lanthanide series  
\* Actinide series

